



**ASU**  
جامعة العلوم التطبيقية الخاصة  
APPLIED SCIENCE PRIVATE UNIVERSITY

AMMAN - JORDAN



المستوى الذهبي

جامعة كل العرب

**Faculty of Arts and Science**  
**Course Syllabus**  
**Semester: First**  
**Academic Year 2020 / 2021**

<b>Course Title:</b>	General Practical Chemistry for Pharmacy Students
<b>Course No.:</b>	1501135
<b>Prerequisite:</b>	-
<b>Concurrent:</b>	-
<b>Department:</b>	Basic Science and Humanities
<b>Coordinator:</b>	Dr. Waed Alahmad

**\* Instructor:**

Name	Office Number	Office Phone	Office Hours	E-mail
Dr. Waed Al Ahmad	221	1414	Announced	<a href="mailto:w_alahmad@asu.edu.jo">w_alahmad@asu.edu.jo</a>
Dr. Ahmad Abu Rayyan	220	1410	At the office	<a href="mailto:a_aburayyan@asu.edu.jo">a_aburayyan@asu.edu.jo</a>

**\* Course Description:**

This course includes selected experiments to illustrate chemical and physical properties, empirical formula, limiting reactant, Le Chatelier's principle, acid base titration and the factor affecting on the rate of reaction, molar mass of volatile liquid.

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رمز النموذج: UF 28 / 1

رقم القرار 24 / 233

تاريخ الاعتماد 2017/9/18



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## جامعة كل العرب

### \* Learning Outcomes:

Upon completion of the course, this module should lead to the following learning outcomes:

#### **A- Knowledge and Understanding (student should):**

Understand the general procedures for conducting various experiments  
know how to collect and organize experimental data  
Identify chemistry laboratory equipment

#### **B- Intellectual Skills:**

The important safety precautions that should be practiced in the laboratory:

- To recognize flammable, toxic, and harmful compounds.
- To practice the procedures for operating common laboratory equipment. (as balances, Bunsen burner, pipette..... etc.)

#### **C- Subject Specific Skills:**

- To emphasize concepts of moles, molar ratio, empirical formula
- To attain a depth of understanding of fundamentals and a reasonable competence in dealing with chemical problems

#### **D- Transferable Skills:**

To test the topics covered, the emphasis on chemical calculations and the mathematical formulation of principles,



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**\* Course Contents and Schedule:**

Semester Weeks	Subject
1st.	Introduction
2nd.	Laboratory Safety and Guidelines
3rd.	Basic Laboratory Operations
4th.	Density Measurement
5th.	Identification of a Compound, Physical Properties
6th.	Determining the Empirical Formula of Magnesium Oxide
7th.	Avogadro's Number
8th.	Limiting Reactant
9th.	Practical Midterm Exam
10th.	Formula of a Hydrate
11th.	Identification of a Compound, Chemical Properties
12th.	Separation of the components of a mixture
13th.	Acid Base Titration
14th.	Revision week
	Final Exam

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## جامعة كل العرب

### \* Teaching Methods:

- Inter active lectures
- Group discussions
- Self reading
- Written assignments

### \* Evaluation:

Midterm Exam	30%
Reports	20%
Final Exam	50%
Total	100%

### \* Written Assignment:

Pre laboratory questions and post laboratory questions for each experiment.

### \* Textbook:

**Manual Book, Editing By Rami Shariah 2017.**

### \* References:

1. Chemistry, Matter and Its Changes; By: Brady, Russell & Holum, 3E, J.W., 2000
2. Chemistry, The Study of Matter and Its Changes by: Brady & Holum 2nd. E. 1996
3. General Chemistry by Zumdal 5<sup>th</sup> E. McGraw Hill. 2000

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## جامعة كل العرب

**Subject Coordinator:**

**Dr. Waed Alahmad**

**Signature:**

**Head of Curriculum Committee**

**Dr. Hussam Miqdad**

**Signature:**

**Department Head**

**Dr. Hussam Miqdad**

**Signature:**

**Dean Faculty**

**Dr. Hadeel Saed**

**Signature:**

**Copy to:**

Department Head. -

Head of Curriculum Committee. -

Course File. -

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